

76. The complex $[\text{Mn}(\text{H}_2\text{O})_6]^{2+}$ has very light pink colour. The best reason for it is
 (a) The complex does not have a charge transfer transition.
 (b) d-d transitions here are orbital forbidden but spin allowed.
 (c) d-d transitions here are orbital allowed but spin forbidden.
 (d) d-d transitions here are both orbital forbidden and spin forbidden.
77. The highest occupied MO in N_2 and O_2^+ respectively are (take x-axis as internuclear axis)
 (a) $\sigma 2p_x, \pi^* 2p_y$ (b) $\pi 2p_y, \pi 2p_z$ (c) $\sigma^* 2p_x, \sigma 2p_x$ (d) $\pi^* 2p_y, \pi^* 2p_z$
78. The correct order of LMCT energies is:
 (a) $\text{MnO}_4^- < \text{CrO}_4^{2-} < \text{VO}_4^{3-}$ (b) $\text{MnO}_4^- > \text{CrO}_4^{2-} > \text{VO}_4^{3-}$
 (c) $\text{MnO}_4^- > \text{CrO}_4^{2-} < \text{VO}_4^{3-}$ (d) $\text{MnO}_4^- < \text{CrO}_4^{2-} > \text{VO}_4^{3-}$
79. Carboxypeptidase contains:
 (a) Zn(II) and hydrolyses CO_2 . (b) Zn(II) and hydrolyses peptide bonds.
 (c) Mg(II) and hydrolyses CO_2 . (d) Mg(II) and hydrolyses peptide bonds.
80. In the EPR spectrum of tetragonal Cu(II) complex, when $g_{\parallel} > g_{\perp} > g_e$ the unpaired electron resides in the orbital.
 (a) d_{xy} (b) $d_{x^2-y^2}$ (c) d_z^2 (d) d_{xz}
81. The oxidative addition and reductive elimination steps are favoured by
 (a) Electron rich metal centres.
 (b) Electron deficient metal centers
 (c) Electron deficient and electron rich metal centers respectively.
 (d) Electron rich and electron deficient metal centers respectively.
82. Identify the order according to increasing stability of the following organometallic compounds, TiMe_4 , $\text{Ti}(\text{CH}_2\text{Ph})_4$, $\text{Ti}(\text{i-Pr})_4$ and TiEt_4 .
 (Me = methyl, Ph = phenyl, i-Pr = isopropyl, Et = ethyl)
 (a) $\text{Ti}(\text{CH}_2\text{Ph})_4 < \text{Ti}(\text{i-Pr})_4 < \text{TiEt}_4 < \text{TiMe}_4$
 (b) $\text{Ti}(\text{CH}_2\text{Ph})_4 < \text{TiEt}_4 < \text{Ti}(\text{i-Pr})_4 < \text{TiMe}_4$
 (c) $\text{Ti}(\text{i-Pr})_4 < \text{TiEt}_4 < \text{TiMe}_4 < \text{Ti}(\text{CH}_2\text{Ph})_4$
 (d) $\text{TiMe}_4 < \text{TiEt}_4 < \text{Ti}(\text{i-Pr})_4 < \text{Ti}(\text{CH}_2\text{Ph})_4$
83. Among the metals, Mn, Fe, Co and Ni, the ones those would react in its native form directly with CO giving metal carbonyl compounds are:
 (a) Co and Mn (b) Mn and Fe (c) Fe and Ni (d) Ni and Co
84. The molecule with highest number of lone-pairs and has a linear shape based on VSEPR theory is:
 (a) CO_2 (b) I_3^- (c) NO_2^- (d) NO_2^+
85. Given, $\text{Ag}^+ + e \rightarrow \text{Ag}, \quad E_0 = 0.50 \text{ V}$
 $\text{Cu}^{2+} + 2e \rightarrow \text{Cu}, \quad E_0 = 0.34 \text{ V}$

A 100 ml solution is 1080 mg with respect to Ag^+ and 635 mg with respect to Cu^{2+} . If 0.1mg Ag^+ left in the solution is considered to be the complete deposition of Ag^+ , the cathode potential, so that no copper is deposited during the process, is:

- (a) 0.16 V (b) 0.84 V (c) 0.31 V (d) -0.16 V

86. In the $\text{H}_2\text{Ru}_6(\text{CO})_{18}$ cluster, containing 8-coordinated Ru centers, the hydrogen atoms are
 (a) Both terminal (b) One terminal and the other bridging
 (c) Both bridging between two Ru centers (d) Both bridging between three Ru centers.
87. In the hydroformylation reaction, the intermediate $\text{CH}_3\text{CH}_2\text{CH}_2\text{Co}(\text{CO})_4$:
 (a) Forms an acyl intermediate $\text{CH}_3\text{CH}_2\text{CH}_2\text{COCO}(\text{CO})_3$
 (b) Forms an adduct with an olefin reactant.
 (c) Reacts with H_2 .
 (d) Eliminates propane.
88. **Statement I:** The sizes of Zr and Hf are similar
Statement II: Size of Hf is affected by lanthanide contraction.
 (a) Statement I and II are correct and II is correct explanation of I.
 (b) Statement I and II are correct but II is not a correct explanation of I.
 (c) Statement I is correct and II is incorrect
 (d) Statements I and II both are incorrect.
89. Consider the compounds, (A) SnF_4 , (B) SnCl_4 and (C) R_3SnCl . The nuclear quadrupole splitting are observed for.
 (a) (A), (B) and (C) (b) (A) and (B) only (c) (B) and (C) only (d) (A) and (C) only
90. Consider two redox pairs
 (1) Cr(II)/Cr(III) (2) Cr(II)/Co(III)
 The rate of acceleration in going from an outer-sphere to an inner-sphere mechanism is lower for (1) relative to (2). Its correct explanation is:
 (a) HOMO/LUMO are σ^* and σ^* respectively. (b) HOMO/LUMO are σ^* and π^* respectively.
 (c) HOMO/LUMO are π^* and σ^* respectively. (d) HOMO/LUMO are π^* and π^* respectively.
91. The correct value of isomer shift (in Mossbauer spectra) and its explanation for Fe(II)-TPP and Fe(III)-TPP respectively from the following are:
 (TPP = tetraphenylporphyrinate)
 (A) 0.52 mms^{-1} (B) 0.45 mms^{-1}
 (C) Increase in s electron density (D) Decrease in s electron density.
 (a) (A) and (D); (B) and (C) (b) (A) and (C); (B) and (C)
 (c) (B) and (D); (A) and (D) (d) (B) and (D); (A) and (C)
92. In IR spectrum of $[\text{Co}(\text{CN})_5\text{H}]^{3-}$ the Co-H stretch is observed at 1840 cm^{-1} . The (Co-D) stretch in $[\text{Co}(\text{CN})_5\text{D}]^{3-}$ will appear at nearly
 (a) 1300 cm^{-1} (b) 1400 cm^{-1} (c) 1500 cm^{-1} (d) 1600 cm^{-1} .
93. For the complexes
 (A) $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ (B) $[\text{Mn}(\text{H}_2\text{O})_6]^{2+}$ (C) $[\text{Cr}(\text{H}_2\text{O})_6]^{3+}$ (D) $[\text{Ti}(\text{H}_2\text{O})_6]^{3+}$, the ideal octahedral geometry will not be observed in
 (a) (A) and (D) (b) (C) and (D) (c) (B) only (d) (D) only
94. Among the following, the number of anhydrides of acids are CO , NO , N_2O , B_2O_3 , N_2O_5 , SO_3 and P_4O_{10} .
 (a) 3 (b) 4 (c) 5 (d) 6

95. For a given nuclear fission reaction of ^{235}U
- $$^{235}_{92}\text{U} + {}^1_0\text{n} \longrightarrow {}^{142}_{56}\text{Ba} + {}^{91}_{36}\text{Kr} + 3{}^1_0\text{n}$$
- the amount of energy (in kJ/mol) released during this process is (given $^{235}\text{U} = 235.0439$ amu, $^{142}\text{Ba} = 141.9164$ amu, $^{91}\text{Kr} = 90.9234$ amu, neutron = 1.00866 amu)
- (a) 3.12×10^{12} (b) 2.8×10^{11} (c) 1.0×10^9 (d) 1.68×10^{10} .
96. The decomposition of gaseous acetaldehyde at T(K) follows second order kinetics. The half-life of this reaction is 400 s when the initial pressure is 250 Torr. What will be the rate constant (in $\text{Torr}^{-1} \text{s}^{-1}$) and half-life (in s) respectively, if the initial pressure of the acetaldehyde is 200 Torr at the same temperature?
- (a) 10^5 and 500 (b) 10^{-5} and 400 (c) 10^{-4} and 400 (d) 10^{-5} and 500
97. For an enzyme catalyzed reaction, a Lineweaver-Burk plot gave the following data:
slope = 40 s
intercept = 4 ($\text{mmol dm}^{-3} \text{s}^{-1}$) $^{-1}$.
If the initial concentration of enzyme is $2.5 \times 10^{-9} \text{ mol dm}^{-3}$, what is the catalytic efficiency (in $\text{dm}^{-3} \text{mol}^{-1} \text{s}^{-1}$) of the reaction?
- (a) 10^5 (b) 10^6 (c) 10^7 (d) 10^4 .
98. A hydrogenic orbital with radial function of the form $r^\alpha \exp[-\beta r]$ and ϕ -part as $\exp[-3i\phi]$ corresponds to
- (a) $n > 4, \ell > 3, m = 3$ (b) $n = 4, \ell = 3, m = -3$
(c) $n = 4, \ell > 3, m = 3$ (d) $n > 4, \ell = 3, m = -3$
99. For an assembly of molecules (molar mass = M) at temperature T, the standard deviation of Maxwell's speed is approximately
- (a) $0.7 \sqrt{\frac{RT}{M}}$ (b) $1.4 \sqrt{\frac{RT}{M}}$ (c) $0.7 \sqrt{\frac{M}{RT}}$ (d) $1.4 \sqrt{\frac{M}{RT}}$
100. The unperturbed energy levels of a system are $\epsilon_0 = 0, \epsilon_1 = 2$ and $\epsilon_2 = 4$. The second order correction to energy for the ground state in presence of the perturbation V for which $V_{10} = 2, V_{20} = 4$ and $V_{12} = 6$ has been found to be
- (a) -6 (b) 0 (c) +6 (d) -8
101. Given the character table of the point group C_{3v} .

	E	$2C_3$	$3\sigma_v$	
A_1	1	1	1	z
A_2	1	1	-1	
E	2	-1	0	(x,y)

Consider the reducible representation, Γ

	E	$2C_3$	$3\sigma_v$
Γ	6	3	0

Its irreducible components are

- (a) $E + 2A_1 + 2A_2$ (b) $2E + A_1 + A_2$ (c) $3A_1 + 3A_2$ (d) $E^2 + 2A_1$
102. Refer to the character table of the point group C_{3v} given above. Find which of the following transition is forbidden
- (a) $a_1 \longleftrightarrow a_1$ (b) $a_1 \longleftrightarrow e$ (c) $a_2 \longleftrightarrow e$ (d) $a_1 \longleftrightarrow a_2$

103. The electronic configuration for gadolinium (Gd) is $[\text{Xe}]4f^7 5d^1 6s^2$, where as that of Gd^{2+} is:
 (a) $[\text{Xe}]4f^5 5d^6 s^2$ (b) $[\text{Xe}]4f^6 6s^2$ (c) $[\text{Xe}]f^6 5d^1 6s^1$ (d) $[\text{Xe}]4f^7 5d^1$
104. The possible J values for 3D term symbol are
 (a) 2 (b) 3 (c) 4 (d) 5
105. The energy levels for cyclobutadiene are $\alpha + 2\beta$, α , α and $\alpha - 2\beta$. The delocalization energy in this molecule is:
 (a) 0 (b) -4β (c) -8β (d) 4α
106. The variation of equilibrium constant (K) of a certain reaction with temperature (T) is
 $\ln k = 3.0 + \frac{2.0 \times 10^4}{T}$ given $R = 8.3 \text{ Jk}^{-1} \text{ mol}^{-1}$, the values of ΔH^0 and ΔS^0 are.
 (a) 166 kJ mol^{-1} and $24.9 \text{ Jk}^{-1} \text{ mol}^{-1}$. (b) 166 kJ mol^{-1} and $-24.9 \text{ Jk}^{-1} \text{ mol}^{-1}$
 (c) -166 kJ mol^{-1} and $24.9 \text{ Jk}^{-1} \text{ mol}^{-1}$ (d) -166 kJ mol^{-1} and $24.9 \text{ Jk}^{-1} \text{ mol}^{-1}$
107. The chemical potential of component 1 in a solution of binary mixture is $\mu_1 = \mu_1^0 + RT \ln p_1$, when p_1 is the partial pressure of component 1 vapour phase. The standard state μ_1^0 is:
 (a) Independent of temperature and pressure
 (b) Depends on temperature and pressure both
 (c) Depends on temperature only
 (d) Depends on pressure only
108. Debye-Hückel screening length (κ^{-1}) is a measure of size of diffuse ion cloud around an ion, provided $\sqrt{\frac{2e^2 N_A}{\epsilon_0 k_B T}} \approx 30 \left(\text{nm} \sqrt{\text{molkg}^{-1}} \right)^{-1}$ at 298K, which of the following values of κ^{-1} is true for a 0.03 molal solution for Na_2SO_4 in water ($\epsilon_r \approx 100$)?
 (a) $\frac{10}{9} \text{ nm}$ (b) $\frac{9}{10} \text{ nm}$ (c) $\frac{10\sqrt{2}}{9} \text{ nm}$ (d) $\frac{9}{10\sqrt{2}} \text{ nm}$
109. If the ratio of composition of oxidized and reduced species in electrochemical cell, is given as $\frac{[\text{O}]}{[\text{R}]} = e^2$, the correct potential difference will be
 (a) $E - E^0 = + \frac{2RT}{nF}$ (b) $E - E^0 = - \frac{2RT}{nF}$
 (c) $E - E^0 = \frac{RT}{nF}$ (d) $E - E^0 = - \frac{RT}{nF}$
110. If the equilibrium constants for the reactions 1 and 2
 (1) $\text{CO}(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightleftharpoons \text{CO}_2(\text{g}) + \text{H}_2(\text{g})$ (2) $\text{CH}_4(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightleftharpoons \text{CO}(\text{g}) + 3\text{H}_2(\text{g})$
 are k_1 and k_2 , the equilibrium constant for the reaction
 $\text{CH}_4(\text{g}) + 2\text{H}_2\text{O}(\text{g}) \rightleftharpoons \text{CO}_2(\text{g}) + 4\text{H}_2(\text{g})$ is:
 (a) $k_1 + k_2$ (b) $k_1 - k_2$ (c) $k_1 k_2$ (d) k_1 / k_2

111. The virial expansion for a real gas can be written in either of the following forms:

$$\frac{P\bar{V}}{RT} = 1 + B_p P + C_p P^2 + \dots$$

$$= 1 + B_v V + C_v V^2 + \dots$$

If $B_v = \alpha B_p$, the value of α would be

- (a) PV/RT (b) RT/PV (c) PV (d) RT

112. A certain system of noninteracting particles has the single-particle partition function $f = A \frac{T^m}{V}$ where A is some constant. The average energy per particle will be

- (a) $m\kappa T$ (b) $A\kappa T$ (c) $\kappa T / m$ (d) $\kappa T / A$

113. Observe the following aqueous solutions of same compound. All the measurements are made at same wavelength and same temperature.

Solution A: The transmittance of 0.1 mol dm^{-3} using 1 cm cell is 0.5 .

Solution B: The optical density 0.5 mol dm^{-3} is measured using 1 mm cell.

Solution C: The transmittance of this solution is 0.1 .

The optical density of these solutions follow the order.

($\log 20 = 1.3010$; $\log 30 = 1.4771$, $\log 50 = 1.6990$)

- (a) $A > B > C$ (b) $B > C > A$ (c) $B > A > C$ (d) $C > A > B$

114. The rotational constant of $^{14}\text{N}_2$ is 2 cm^{-1} . The wave number of incident radiation in a Raman spectrometer is 20487 cm^{-1} . What is the wave number of first scattered Stokes line (in cm^{-1}) of $^{14}\text{N}_2$?

- (a) 20479 (b) 20475 (c) 20499 (d) 20495

115. For a certain particle encountering a barrier, the tunneling probability is approximately e^{-10} . If the mass is halved and width of the barrier (rectangular) doubled, approximate value of the tunneling probability will be

- (a) $e^{-10/\sqrt{2}}$ (b) $e^{-10\sqrt{2}}$ (c) $e^{-20\sqrt{2}}$ (d) e^{-10}

116. An operator A is defined as $A = -\frac{d}{dx} + x$. Which one of the following statements is true?

- (a) A is a Hermitian operator. (b) A^\dagger is an antihermitian operator.
 (c) Both AA^\dagger and $A^\dagger A$ are Hermitian. (d) AA^\dagger is Hermitian, but $A^\dagger A$ is antihermitian.

117. Isotherm which has fractional coverage, linearly, dependent on pressure at low pressures but almost independent at high pressure is called

- (a) BET isotherm (b) Langmuir isotherm
 (c) Freundlich isotherm (d) Temkin isotherm

118. A one-dimensional crystal of lattice dimension 'a' is metallic. If the structure is distorted in such a way that the lattice dimension is enhanced to '2a'.

- (a) The electronic structure remains unchanged.
 (b) The width of conduction band decreases and a band gap is generated.
 (c) The width of conduction band increases
 (d) The width of the conduction band remains unchanged.

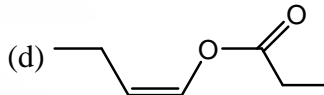
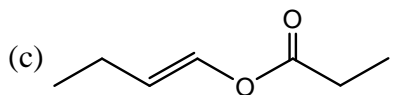
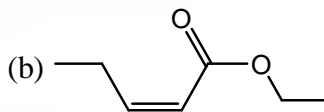
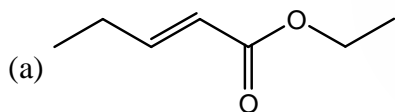
119. For a H_2 molecule, the ground state wavefunction is $\psi(1,2) = \phi(1,2)\sigma(1,2)$ where ϕ refers to the space part and σ to the spin part. Given that $\phi(1,2) = \phi(2,1)$, the form of $\sigma(1,2)$ would be
- (a) $\alpha(1)\beta(2)$ (b) $\alpha(2)\beta(1)$
 (c) $\alpha(1)\beta(2) - \alpha(2)\beta(1)$ (d) $\alpha(1)\beta(2) + \alpha(2)\beta(1)$

120. There are several types of mean molar masses for polymer and they are dependent on experimental methods like:

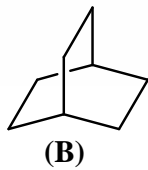
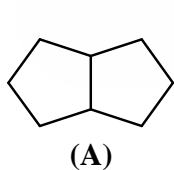
- (1) Osmometry (2) Light scattering (3) Sedimentation.

Correct relation between mean molar masses and experimental method is:

- (a) $\bar{M}_n \Leftrightarrow (3), \bar{M}_w \Leftrightarrow (2), \bar{M}_z \Leftrightarrow (1)$ (b) $\bar{M}_n \Leftrightarrow (2), \bar{M}_w \Leftrightarrow (3), \bar{M}_z \Leftrightarrow (1)$
 (c) $\bar{M}_n \Leftrightarrow (1), \bar{M}_w \Leftrightarrow (2), \bar{M}_z \Leftrightarrow (3)$ (d) $\bar{M}_n \Leftrightarrow (1), \bar{M}_w \Leftrightarrow (3), \bar{M}_z \Leftrightarrow (2)$
121. An organic compound ($C_7H_{12}O_2$) exhibited the following data in the 1H NMR spectrum.
 δ 7.10(1H, d t, $J = 16$ and 7.2 Hz), 5.90(1H, d t, $J = 16$ and 2 Hz),
 4.1(2H, q, $J = 7.2$ Hz), 2.10(2H, m), 1.25(3H, t, $J = 7.2$ Hz),
 0.90(3H, t, $J = 7.2$ Hz) ppm.
 The compound, among the choices given below, is:

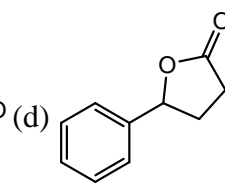
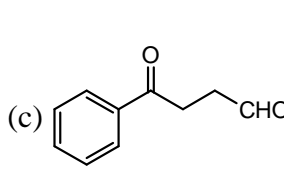
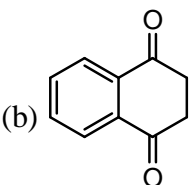
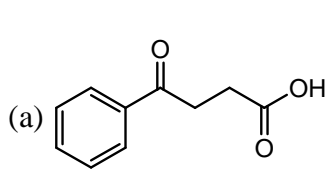
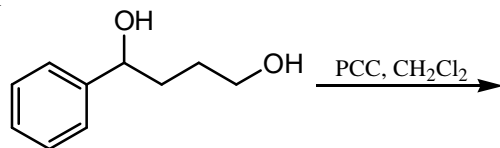


122. In the broadband decoupled ^{13}C NMR spectrum, the number of signals appearing for the bicyclooctane A-C, respectively, are

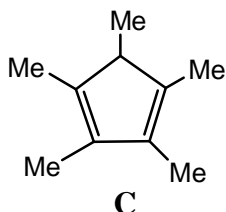
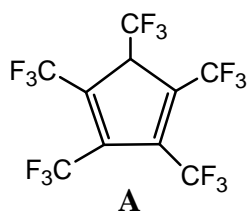


- (a) Five, four and eight (b) Three, two and five
 (c) Five, four and five (d) Three, two and eight
123. In the mass spectrum of dichlorobenzene the ratio of the peaks at m/z 146, 148 and 150, is:
 (a) 1:1:1 (b) 3:3:1 (c) 1:2:1 (d) 9:6:1

124. The major compound X formed in the following reaction exhibited a strong absorption at ν_{max} 1765 cm^{-1} in the IR spectrum. The structure of X is:

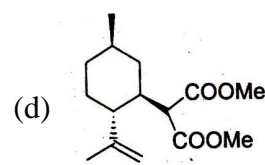
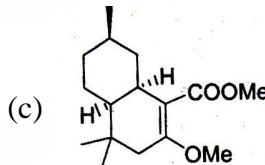
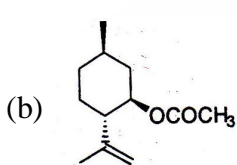
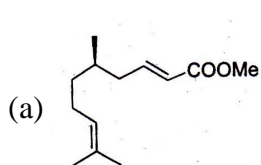
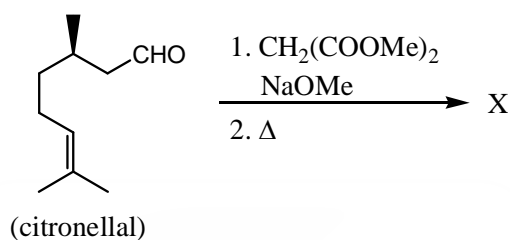


125. The correct order of acidity of the following compound A–C is:

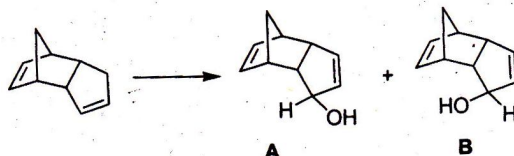


- (a) $B > C > A$ (b) $C > B > A$ (c) $A > C > B$ (d) $A > B > C$

126. The major product formed in the reaction sequence is:

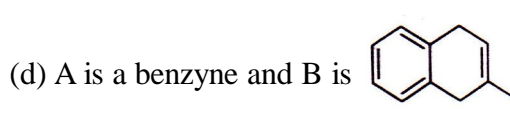
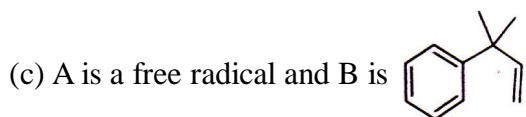
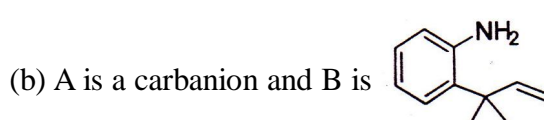
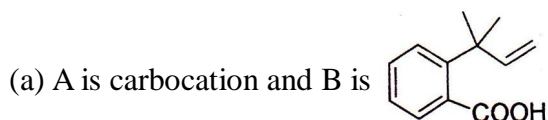
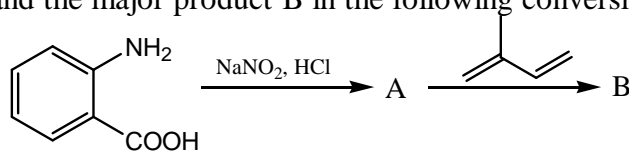


127. For the following allylic oxidation reaction, the appropriate statement, among the choices given below, is:

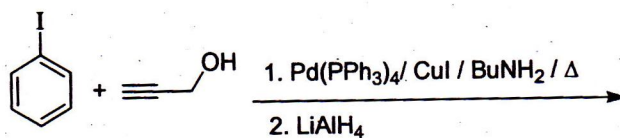


- (a) Suitable reagent is KMnO_4 and the major product is A.
 (b) Suitable reagent is KMnO_4 and the major product is B.
 (c) Suitable reagent
 (d) Suitable reagent is SeO_2 and the major product is B.

128. The intermediate A and the major product B in the following conversion are

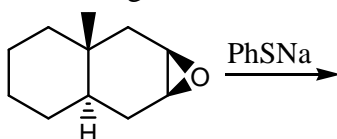


129. The major product formed in the following reaction is:



- (a) C=CCOc1ccccc1 (b) OCC=CCc1ccc(I)cc1
 (c) OCC=CCc1ccccc1 (d) OCC=CCc1ccccc1

130. The major product formed in the following reaction is:

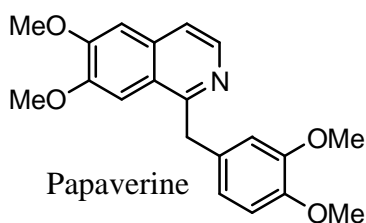


- (a) C12CCC3C1CC2(O)C3S(=O)(=O)c4ccccc4 (b) C12CCC3C1CC2(O)C3C4=CC=CC=C4
 (c) C12CCC3C1CC2(S(=O)(=O)c4ccccc4)C3O (d) C12CCC3C1CC2(O)C3C4=CC=CC=C4

131. The major product formed in the reaction of glucose with benzaldehyde and p-TSA is:

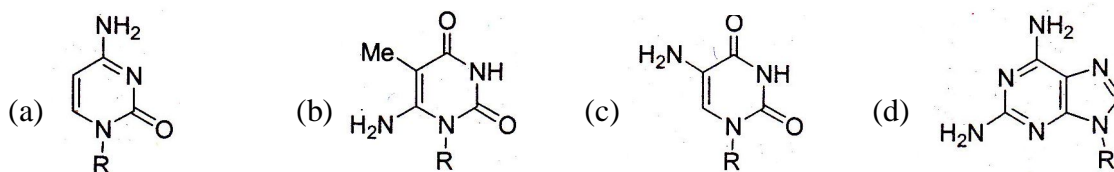
- (a) O[C@H]1O[C@@H](OC2=CC=CC=C2)[C@H](O)[C@@H](O)[C@H]1O (b) O[C@H]1O[C@@H](OC2=CC=CC=C2)[C@H](O)[C@@H](OC3=CC=CC=C3)[C@H]1O
 (c) O[C@H]1O[C@@H](OC2=CC=CC=C2)[C@H](O)[C@@H](O)[C@H]1O (d) O[C@H]1O[C@@H](OC2=CC=CC=C2)[C@H](O)[C@@H](O)[C@H]1O

132. Papaverine on oxidation with potassium permanganate gives a ketone, which on fusion with potassium hydroxide gives

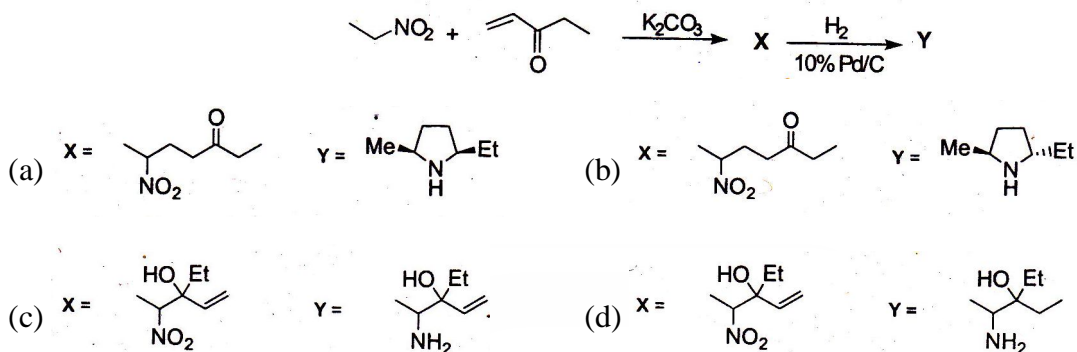


- (a) COc1cc(C(=O)O[K])ccc1OC (b) COc1cc(C(=O)O[K])cc(OC)c1 (c) COc1ccc2nc3cc(OC)ccc3cc21 (d) COc1ccc(C(=O)O[K])cn1

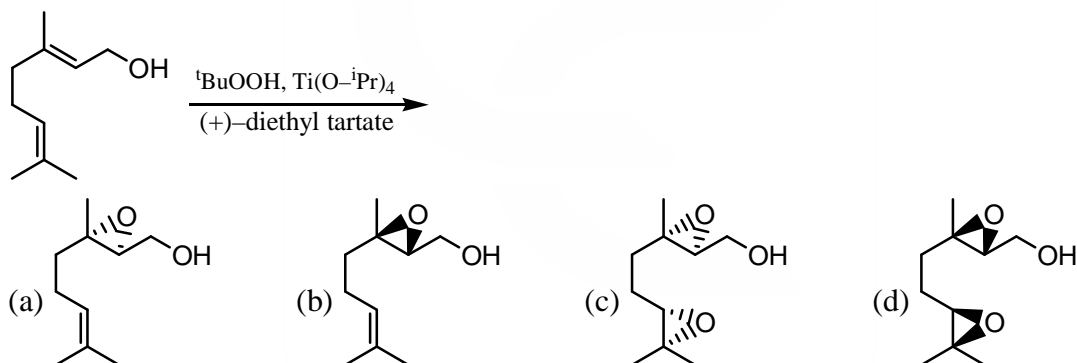
133. The major product formed on nitration ($\text{HNO}_3/\text{H}_2\text{SO}_4$) of uridine followed by reduction with tin and HCl is:



134. In the following reaction sequence, the correct structures for the major products X and Y are

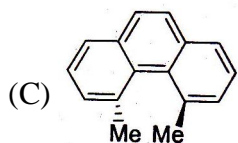
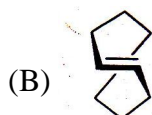
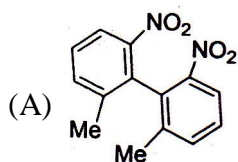


135. The major product formed in the following reaction is:



136. Match the following:

Column I



Column II

(i) Chiral centre

(ii) Chiral axis

(iii) Chiral plane

(iv) Helical chirality

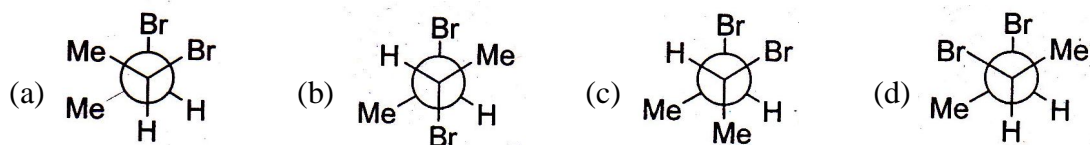
(a) (A)-iii (B)-ii (C)-iv

(b) (A)-iv (B)-iii (C)-ii

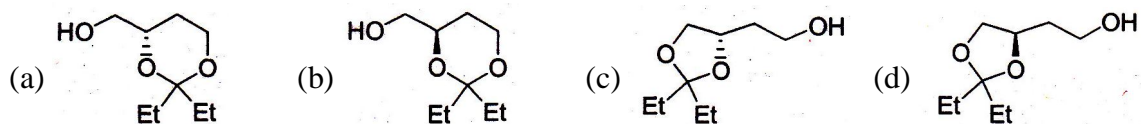
(c) (A)-ii (B)-iv (C)-iii

(d) (A)-ii (B)-iii (C)-iv

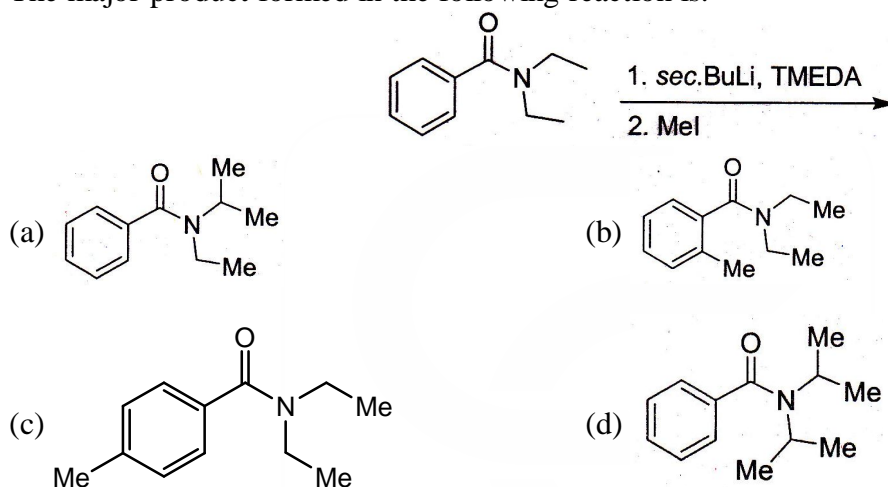
137. The gauche interaction values for Me/Me, Me/Br and Br/Br are 3.3, 0.8 and 3.0 kJ/mol, respectively. Among the following, the most stable conformation of 2, 3-dibromobutane is:



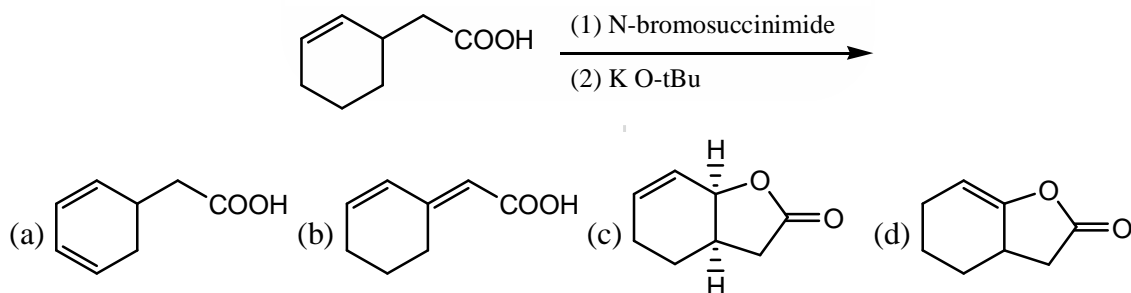
138. The major product formed in the reaction of (S)-1, 2, 4-butanetriol with 3-pentanone in the presence of a catalytic amount of p-TSA is:



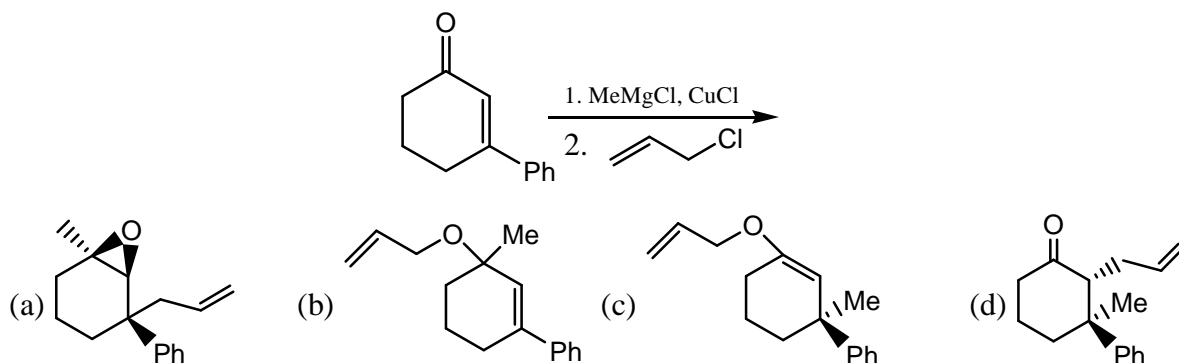
139. The major product formed in the following reaction is:



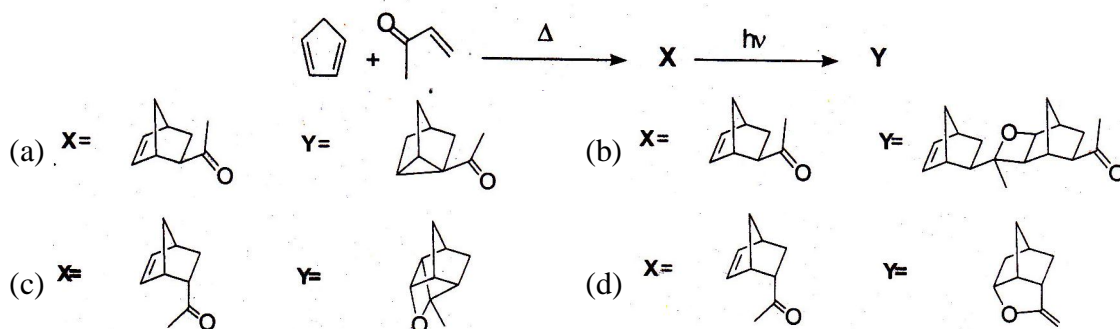
140. The major product formed in the following transformation is:



141. The major product formed in the following transformation is:



142. The structures of the major products X and Y in the following transformation are

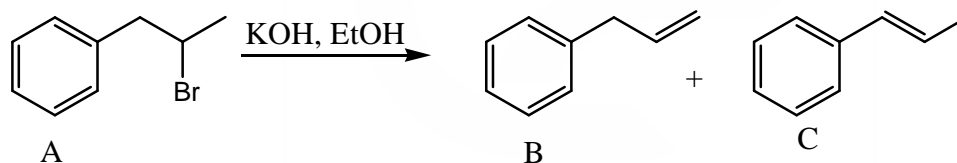


143. Match the following:

Column I	Column II
A. Pyrrole	i. Pictet-Spengler
B. 1, 4-dihydropyridine	ii. Chichibabin
C. Isoquinoline	iii. Paal-Knorr
	iv. Hantzsch

- (a) (A)-i (B)-ii (C)-iii (b) (A)-ii (B)-iii (C)-iv
 (c) (A)-iv (B)-i (C)-ii (d) (A)-iii (B)-iv (C)-i

144. Consider the following reaction:



In an experiment, 1.99 g of bromide A on reaction with ethanolic potassium hydroxide gave 1.062 g of a mixture of the olefins B and C. If the ratio of olefins B:C formed is 2:1, the yields for their formation, respectively, are

- (a) 60 and 30% (b) 50 and 25% (c) 66 and 33% (d) 54 and 27%

145. An organic compound A ($C_8H_{16}O_2$) on treatment with an excess of methylmagnesium chloride generated two alcohols B and C, whereas reaction of A with lithium aluminium hydride generated only a single alcohol C. Compound B on treatment with an acid yielded an olefin (C_6H_{12}), which exhibited only a singlet at δ 1.6 ppm in the 1H NMR spectrum. The compound A is:

